**Physical Science Chemical Reactions Review**

**1. Identify the reactants and products in the following chemical equation:**

**Cd(NO3)2 *(aq)* + H2S *(g)* 🡪 CdS *(s)* + 2HNO3 *(aq)***

**2. What is the Law of Conservation of Mass?**

**3. Balance the Equation:**

 **Fe *(s)* + Cl2 *(g)* 🡪 FeCl3 *(s)***

**4. What coefficient is assumed if no coefficient is written before a formula in a chemical equation?**

**5. Classify the following reactions:**

**CaO + H2O 🡪 Ca(OH)2**

**NH4NO3  🡪 N2O + 2H2O**

**6. Wood Burning and the explosion of dynamite are examples of an \_\_\_\_\_\_\_\_\_\_\_\_\_ reaction.**

**7. An example of an \_\_\_\_\_\_\_\_\_\_\_\_\_ reaction is a cold pack of ammonium nitrate crystals and water.**

**8. What are 2 examples of Energy that can be released in an exothermic reaction?**

**9. Because some reactions proceed too slowly, sometimes a \_\_\_\_\_\_\_\_\_\_\_\_ is needed to speed it up.**

**10. What type of compound is the food additive BHA?**

**11. What do the symbols *(s)*, *(aq)*, *(g)* and *(l)* mean when written next to a formula of a compound in a chemical reaction?**

**12. Magnesium reacts with Oxygen to form Magnesium Oxide. Please write the balanced chemical equation.**

**13. What type of compound represents a precipitate in a chemical equation?**

**14. A metal will replace any metal that is \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_. (Pg 643)**